

## Extension Questions

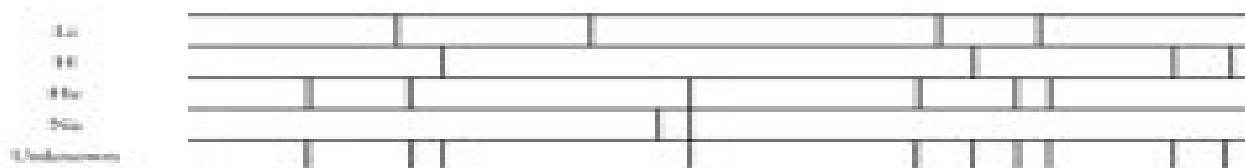
24. The hydrogen spectral lines in Model 2 are only the wavelengths of light that are in the visible range and therefore "seen" by the naked eye. However, many other wavelengths can be detected with special equipment.
- Propose a hydrogen electron transition that involves light with a wavelength in the ultraviolet (UV) range (10–400 nm).

$$n=6 \rightarrow n=1$$

- Propose a hydrogen electron transition that involves light with a wavelength in the infrared (IR) range (1000–106 nm).

$$n=6 \rightarrow n=5$$

25. Below are diagrams for the bright line spectra of four elements and the spectrum of a mixture of unknown gases.



- Which element(s) are not present in the Unknown?  
*Elements that are metals aren't in the unknown.*
  - Which element(s) are in the Unknown?  
*The noble gases are in the unknown.*
26. Model 2 shows the emission spectra for hydrogen and boron. Scientists can also record the absorption spectra for elements. Propose how this might be done, and what the absorption spectra of hydrogen and boron would look like.

*The light spectrum of these elements could be displayed by projecting light through a prism. The spectrum of these would be more complex and have more different colors than the spectrum of hydrogen.*

# Electron Energy And Light Extension Questions Pogil Answers

**David K. Adams**



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Electron Energy in Solids Joseph Callaway,1958      **Note on the Electron Energy Spectrum in the Inner Van Allen Belt** Guido Pizzella,C. D. Laughlin,Brian J. O'Brien,1962      **Electron energy levels in crystalline solids** David K. Adams,1952      **Electron energy determinations** R. L. Graham,1960

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