

UNIT 2

ELECTROCHEMISTRY CONCEPTS

Points to Remember

Electrochemistry may be defined as the branch of chemistry which deals with the quantitative study of inter-relationship between chemical energy and electrical energy and inter-conversion of one form into another relationships between electrical energy taking place in redox reactions.

A cell is of two types :

- I. Galvanic cell
- II. Electrolytic cell

In Galvanic cell, the chemical energy of a spontaneous redox reaction is converted into electrical work.

In Electrolytic cell, electrical energy is used to carry out a non-spontaneous redox reaction.

1. **Conductivity (k) :**

$$k = \frac{1}{\rho} = \frac{1}{R} \times \frac{l}{A}$$

where R is Resistance, l/A = cell constant (G^*) and ρ is resistivity.

2. **Relation between k and Λ_m**

$$\Lambda_m = \frac{1000 \times k}{C}$$

where Λ_m is molar conductivity, k is conductivity and C is molar concentration.

Kohlrausch's law :

(a) In general, if an electrolyte on dissociation give γ_+ cations and γ_- anions, then its limiting molar conductivity (Λ_m°) is given by

$$\Lambda_m^\circ = \nu_+ \lambda_+^\circ + \nu_- \lambda_-^\circ$$

Here, λ_+° and λ_-° are the limiting molar conductivities of cation and anion respectively and ν_+ and ν_- are the number of cations and anions furnished by one formula unit of the electrolyte.

Electrochemistry Practical Manual

Joseph Lancaster Budd



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